

# Tests, On Your Owns, and Practice Problems

## How did you get $K_2SO_4$ as a product on Module 9 practice problem 10 in chemistry?

The reason you do not know how the "2" got into the chemical formula for the salt is that you have forgotten what you learned in Module #5 (pp. 175-180). When  $H_2SO_4$  gives up its two  $H^{1+}$  ions, the only thing left is  $SO_4^{2-}$ , the polyatomic sulfate ion. The K is in group 1, so in ionic compounds, it takes on the charge  $1+$ .

To determine the chemical formula of the ionic compound between  $SO_4^{2-}$  and  $K^{1+}$ , you switch the charges and drop the signs. Thus, you get  $K_2SO_4$ .

You might want to review those module #5 pages, because you must remember those concepts to understand module #9.

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